

$$70 \text{ g HCl: } \frac{70}{1+35.5} = 1.9 \text{ mol}$$

$$M = \frac{n}{V}$$

$$\text{En 2 L de H}_2\text{O: } M = \frac{1.9}{2} = 0.95 \text{ M}$$

$$100 \text{ g C}_6\text{H}_{12}\text{O}_6: \frac{100}{12 \cdot 6 + 1 \cdot 12 + 16 \cdot 6} = 0.56 \text{ mol}$$

$$\text{En 5 L H}_2\text{O: } M = \frac{0.56}{5} = 0.11 \text{ M}$$

$$M = \frac{n}{V}$$

$$3 = \frac{n}{20} \Rightarrow 3 \cdot 20 = n \Rightarrow n = 60 \text{ mol NaCl}$$

$$60 \cdot (23 + 35.5) = 3510 \text{ g NaCl}$$

$$2 = \frac{n}{15} \Rightarrow 2 \cdot 15 = n \Rightarrow n = 30 \text{ mol C}_6\text{H}_{12}\text{O}_6$$